



## INTERNATIONAL APPLICATION PUBLISHED UNDER THE PATENT COOPERATION TREATY (PCT)

(51) International Patent Classification 6: <b>B01J 20/28, 20/10</b>	<b>A1</b>	(11) International Publication Number: <b>WO 98/31461</b> (43) International Publication Date: <b>23 July 1998 (23.07.98)</b>
<p>(21) International Application Number: <b>PCT/US98/00533</b></p> <p>(22) International Filing Date: <b>7 January 1998 (07.01.98)</b></p> <p>(30) Priority Data: <b>08/786,600</b>      <b>21 January 1997 (21.01.97)</b>      <b>US</b></p> <p>(71) Applicant: <b>W.R. GRACE &amp; CO.-CONN. [US/US]; 1114 Avenue of the Americas, New York, NY 10036 (US).</b></p> <p>(72) Inventors: <b>PRYOR, James, Neil; 3253 Denmark Drive, West Friendship, MD 21794 (US). CRUMP, Linda, Lee; 306 Severn Road, Crownsville, MD 21032 (US).</b></p> <p>(74) Agent: <b>CROSS, Charles, A.; W.R. Grace &amp; Co.-Conn., 7500 Grace Drive, Columbia, MD 21044 (US).</b></p>		<p>(81) Designated States: <b>AL, AM, AT, AU, AZ, BA, BB, BG, BR, BY, CA, CH, CN, CU, CZ, DE, DK, EE, ES, FI, GB, GE, GH, HU, IL, IS, JP, KE, KG, KP, KR, KZ, LC, LK, LR, LS, LT, LU, LV, MD, MG, MK, MN, MW, MX, NO, NZ, PL, PT, RO, RU, SD, SE, SG, SI, SK, SL, TJ, TM, TR, TT, UA, UG, UZ, VN, YU, ZW, ARIPO patent (GH, GM, KE, LS, MW, SD, SZ, UG, ZW), Eurasian patent (AM, AZ, BY, KG, KZ, MD, RU, TJ, TM), European patent (AT, BE, CH, DE, DK, ES, FI, FR, GB, GR, IE, IT, LU, MC, NL, PT, SE), OAPI patent (BF, BJ, CF, CG, CI, CM, GA, GN, ML, MR, NE, SN, TD, TG).</b></p> <p><b>Published</b></p> <p><i>With international search report.</i></p> <p><i>Before the expiration of the time limit for amending the claims and to be republished in the event of the receipt of amendments.</i></p>
<p>(54) Title: <b>SILICA ADSORBENT ON MAGNETIC SUBSTRATE</b></p> <p>(57) Abstract</p> <p>Adsorbent particles comprising superparamagnetic and/or low Curie temperature transition metal-containing cores surrounded by a hydrous siliceous oxide coating can be formed by an aqueous process wherein the core is precipitated from an aqueous solution and a siliceous oxide coating is deposited thereon while complete drying of the core is avoided until after the siliceous oxide is deposited. The resulting siliceous adsorbents exhibit strong superparamagnetic and/or low Curie temperature magnetic properties with low transition metal leachability.</p>		

**FOR THE PURPOSES OF INFORMATION ONLY**

Codes used to identify States party to the PCT on the front pages of pamphlets publishing international applications under the PCT.

AL	Albania	ES	Spain	LS	Lesotho	SI	Slovenia
AM	Armenia	FI	Finland	LT	Lithuania	SK	Slovakia
AT	Austria	FR	France	LU	Luxembourg	SN	Senegal
AU	Australia	GA	Gabon	LV	Latvia	SE	Sweden
AZ	Azerbaijan	GB	United Kingdom	MC	Monaco	TD	Chad
BA	Bosnia and Herzegovina	GE	Georgia	MD	Republic of Moldova	TG	Togo
BB	Barbados	GH	Ghana	MG	Madagascar	TJ	Tajikistan
BE	Belgium	CN	China	ME	The former Yugoslav Republic of Macedonia	TM	Turkmenistan
BF	Burkina Faso	GR	Greece	NL	Netherlands	TR	Turkey
BG	Bulgaria	HU	Hungary	MM	Myanmar	TT	Trinidad and Tobago
BJ	Benin	IE	Ireland	MN	Mongolia	UA	Ukraine
BR	Brazil	IL	Israel	MR	Mauritania	UG	Uganda
BY	Belarus	IS	Iceland	MW	Malawi	US	United States of America
CA	Canada	IT	Italy	MX	Mexico	UZ	Uzbekistan
CF	Central African Republic	JP	Japan	NE	Niger	VN	Viet Nam
CG	Congo	KE	Kenya	NL	Netherlands	YU	Yugoslavia
CH	Switzerland	KG	Kyrgyzstan	NO	Norway	ZW	Zimbabwe
CI	Côte d'Ivoire	KP	Democratic People's Republic of Korea	PL	Poland		
CM	Cameroon	KR	Republic of Korea	PT	Portugal		
CN	China	KZ	Kazakhstan	RO	Romania		
CQ	Cuba	LC	Saint Lucia	RU	Russian Federation		
CR	Costa Rica	LI	Liechtenstein	SD	Sudan		
DE	Germany	LK	Sri Lanka	SE	Sweden		
DK	Denmark	LR	Liberia	SG	Singapore		
EE	Estonia						

## **Silica Adsorbent on Magnetic Substrate**

### **Cross-Reference to Related Applications**

Reference is made to concurrently filed U.S. Patent Application Serial No.  
5 08/786,601, entitled "Methods of Isolating Biological Target Materials Using Silica  
Magnetic Particles", the disclosure of which is incorporated herein by reference.

### **Background of the Invention**

Adsorbent materials having magnetic properties are known. Often magnetic  
10 properties such as ferromagnetism and superparamagnetism are desired to enable  
use of magnetic fields in various processes. Superparamagnetism is of special  
interest in many chemical applications where it is desired to collect and redisperse  
adsorbent particles. The particles of the prior art generally comprise a  
superparamagnetic material (e.g. magnetite) which is enveloped in or treated with  
15 some diverse material.

For many applications, the materials of the prior art have been considered  
adequate, however for certain applications, the materials of the prior art are  
inadequate due to their vulnerability to leaching in acidic environments or their lack of  
sufficient magnetic performance. In many applications such as processing of edible  
20 substances or sensitive biological materials (e.g. nucleic acid purification such as  
described in published patent application WO95/06652), these deficiencies result in  
poor performance or simply prohibit use of these materials.

Thus, there is a need for improved coated superparamagnetic particles,  
especially particles useful as superparamagnetic adsorbents. Further, there is a  
25 need for improved methods of making such adsorbents.

### **Summary of the Invention**

The invention provides improved siliceous oxide-coated magnetic particles  
having a high resistance to leaching of the magnetic material on exposure to  
30 aqueous acidic environments while also possessing a hydrous siliceous oxide  
adsorptive surface and excellent magnetic response.

In one aspect, the invention encompasses particles having a superparamagnetic or low Curie Temperature core surrounded by a hydrous siliceous oxide coating wherein the particles exhibit little or no transition metal leaching on exposure to acidic media. The particles preferably exhibit low metal leaching even after sonication (i.e. ultrasonic treatment).

In another aspect, the invention encompasses methods for forming siliceous coatings on superparamagnetic or low Curie Temperature cores wherein the method comprises forming an aqueous dispersion of the cores and depositing a siliceous oxide coating on said dispersed cores wherein said deposition process has an end pH of about 9 or less. Preferably, the aqueous dispersion of the cores is achieved by precipitation of the cores (crystals) in an aqueous medium and maintaining the cores in contact with an aqueous medium from the moment of their precipitation through the deposition of the silica coating thereon.

In another aspect, the invention encompasses methods for adsorbing substances wherein particles having a superparamagnetic or low Curie Temperature core surrounded by a hydrous siliceous oxide coating are used to adsorb a substance from a system and the adsorbent is subsequently removed from the system by application of a magnetic field.

The core material is preferably a superparamagnetic material such as magnetite.

#### Detailed Description of the Invention

The invention provides improved particles comprising siliceous oxide-coated superparamagnetic or low Curie Temperature cores having a high resistance to leaching of transition metal on exposure to aqueous acidic environments while also possessing a hydrous siliceous oxide adsorptive surface (i.e. a surface characterized by the presence of silanol groups) and excellent magnetic response. The particles of the invention may comprise agglomerates of these siliceous oxide-coated cores.

The particles of the invention are especially characterized by their (1) high resistance to leaching of metal from the magnetic core, especially on exposure to acidic environments, (2) excellent magnetic performance (e.g. their ability to be separated from a liquid by application of a magnetic field and their dispersibility in the

absence of a magnetic field), and (3) excellent adsorption/desorption behavior, especially for adsorption/desorption of genetic biological material from a lysate.

For the purpose of describing this invention, the "core" refers to all the material which is surrounded by the siliceous oxide coating. The core of the  
5 particles of the invention may be any material which is amenable to the silica coating process which material exhibits superparamagnetic performance (including materials which exhibit a relatively low remanent magnetism) and/or has a low Curie  
Temperature. Preferably, the core comprises a superparamagnetic material having a remanent magnetism of less than about 10 emu/g, more preferably less than about  
10 2 emu/g, most preferably 0 to 1 emu/g. Where a low Curie Temperature magnetic material is used, that material preferably has a Curie Temperature below about  
100°C, more preferably between about -50°C and 100°C, most preferably between about 0°C and 90°C. The core materials are generally characterized by the  
presence of one or more transition metals. Of the superparamagnetic materials,  
15 metal oxides containing group VIII transition metal are preferred, magnetite iron oxide being most preferred. The core preferably consists essentially of iron magnetite.

In general, superparamagnetic materials are preferred over low Curie  
Temperature magnetic materials since they do not require a change in temperature  
20 to cause loss of magnetism on removal of an external field. Particles of the superparamagnetic materials may be easily redispersed on removal of any external magnetic field due to their low or non-existent remanent magnetism whereas the low Curie Temperature materials require elevation of the temperature to a level above the Curie Temperature to permit easy redispersion of the particles on removal of any  
25 external magnetic field.

The magnetic behavior of many materials often varies with the physical size and state of the material. Thus, materials may exhibit differing magnetic performance depending on their crystal size, their temperature or their physical surroundings (e.g. if they are embedded in a matrix). For magnetite and most other  
30 superparamagnetic materials, the material should have a crystal size (diameter) less than about 1000Å, more preferably about 600Å or less. In some cases, it may be possible to form the core from a combination of the above-mentioned magnetic materials. Alternatively, it may be possible to dilute the composition of the core with

adsorption and/or desorption depending on the nature of the porosity and the material to be adsorbed and/or desorbed. In cases where both adsorption and subsequent desorption are desired, the pores of the siliceous oxide coating are preferably such that they provide additional surface for adsorption, but do not significantly prevent desorption. In this context, pores of some intermediate size are generally preferred depending on the size of the materials to be adsorbed/desorbed. Such pores will be large enough such that the desired adsorbed species is not excessively inhibited from desorption at a later point in the process. In most circumstances, the external surface of the siliceous oxide coating is also effective to provide adsorption. Thus, accessible internal surface (i.e. porosity) may not be required to achieve a usable product.

In general, the siliceous oxide coating preferably has a total pore volume measured by nitrogen BET method of at least about 0.2 ml/g, more preferably about 0.5 to 1.5 ml/g based on the total mass of the particles. Of the total pore volume measured by nitrogen BET, preferably at least about 50% of the pore volume is contained in pores having a diameter of 600 Å or greater, more preferably at least about 60%, most preferably about 70 to 85%. The siliceous oxide coating preferably has a nitrogen BET surface area, based on the total mass of the particles, of at least 5 m<sup>2</sup>/g, more preferably at least about 30 m<sup>2</sup>/g, most preferably about 40-500 m<sup>2</sup>/g.

As noted above, the particles of the invention are may comprise an agglomeration of siliceous oxide-coated superparamagnetic or low Curie Temperature crystal cores. The agglomerates themselves may have additional siliceous oxide coating on their accessible surfaces. While the relative amounts of the core material and the siliceous oxide coating may be varied considerably, the particles of the invention preferably comprise at least 30 wt. % of the superparamagnetic or low Curie Temperature core material based on the total particle composition, more preferably at least 50 wt. %, most preferably about 60 - 80 wt. %. In general, higher amounts of core material will provide stronger response to an applied magnetic field. When the particles are present in a liquid while the magnetic field is applied, the drag characteristics (resistance to particle movement in the fluid associated with the particle-fluid interface and the particle morphology) of the particles themselves will also affect the strength and speed of response (i.e. particle movement) to the applied field.

The particles of the invention may be made in various sizes. Smaller particles provide more surface area (on a weight basis) for adsorption, but they also tend to exhibit a higher amount of drag. The particles of the invention preferably have a median particle size of about 1-15  $\mu\text{m}$ , more preferably about 3-10  $\mu\text{m}$ , most preferably about 4-7  $\mu\text{m}$ . The particle size distribution may also be varied. In general, a relatively narrow monodal particle size distribution is preferred. The particle size distribution is preferably such that about 80 wt.% of the particles are within a 10  $\mu\text{m}$  range about the median particle size, more preferably within an 8  $\mu\text{m}$  range, most preferably within a 6  $\mu\text{m}$  range.

10       The particles of the invention are generally useful for all applications where siliceous oxide adsorbents are used. Thus, the particles of the invention may be used as selective adsorbents in liquids, as desiccants, odor adsorbents, volatile organic adsorbents, etc. The particles are especially useful for selective adsorption of macromolecular constituents (especially organics) from liquids. The particles are especially useful in mild environments and in highly acidic environments.

15       In processes involving adsorption from liquids, the particles can be easily dispersed in the liquid in the absence of any applied magnetic field. Once the desired adsorption has taken place, the particles may be removed from the liquid by conventional techniques such as centrifugation or filtration, however, the particles of the invention are especially suited for separation from the liquid by application of a magnetic field across the liquid. A further advantage of the particles of the invention is that they can be easily redispersed on removal of the magnetic field. For low Curie Temperature materials, the application of the magnetic field should take place below the relevant Curie Temperature. On removal of the applied magnetic field, the low Curie Temperature material would have to be heated to a point above the Curie Temperature in order to reduce the remanent magnetism and allow redispersion.

20       The particles of the invention are preferably further characterized by their performance in these various applications. Specifically, the particles of the invention are preferably characterized by excellent resistance to leaching of metals from the core on exposure of the particles to acidic environments. The particles of the invention are preferably characterized by excellent response to external magnetic fields as demonstrated by the separation rate of the particles on application of a

the invention preferably meet or exceed the adsorption/desorption performance described in U.S. Patent Application Serial No. \_\_\_\_\_, entitled "Methods of Isolating Biological Target Materials Using Silica Magnetic Particles" filed on the same day as the present application.

5           The particles of the invention may be made by a variety of methods. These methods generally involve the formation of an aqueous slurry of dispersed core particles having the desired magnetic properties and deposition of a siliceous coating onto the surface of the core particles wherein the slurry pH at the end of the deposition is about 9 or less.

10           The core particles may be obtained from commercial sources or they may be synthesized by techniques known in the art for the desired core material. The core particles are preferably thoroughly dispersed in an aqueous medium prior to formation of the siliceous oxide coating thereon. Most preferably, the cores are synthesized via precipitation in an aqueous medium and are maintained in contact  
15 with an aqueous medium from the moment of their precipitation through the deposition of the silica coating thereon.

For magnetite, the crystals are preferably prepared by an aqueous precipitation process wherein an aqueous solution of  $\text{FeCl}_2$  (e.g. about 2.5-4 wt.%) and  $\text{FeCl}_3$  (e.g. about 1-3 wt.%) is formed. The ratio of  $\text{Fe(II)}$  to  $\text{Fe(III)}$  may be  
20 varied via the relative proportions of the iron chlorides to affect the ratio of iron oxides in the final product. In general, the  $\text{Fe(II)}$  to  $\text{Fe(III)}$  ratio is about 1. Under strong agitation, a base such as ammonium hydroxide (preferably as a 14 wt.%  $\text{NH}_4\text{OH}$  aqueous solution) is then rapidly added to the solution at ambient temperature to raise the pH to about 8-10, more preferably about 8.5, whereby  
25 microcrystalline magnetite is precipitated. The precipitate slurry is then settled, decanted, and reslurried with water to remove excess ammonium and chloride ions. The settling, decanting and washing steps may be repeated to reduce the level of undesired ions. If desired, centrifugation may be used in place of the settling and decantation steps to accelerate the overall process. Drying of the precipitate prior to  
30 formation of the silica coating is preferably avoided.

As noted above, it may be possible to form one or more intermediate layers of a diverse non-siliceous oxide material on the surface of the superparamagnetic or low Curie Temperature material. For example, the surface of the core may be

subjected to a mildly oxidizing environment to produce a surface region of slightly higher oxygen content. Most preferably, all intermediate layers are avoided such that the siliceous oxide coating directly contacts the surface of the superparamagnetic or low Curie Temperature material.

- 5 For deposition of the siliceous oxide coating, the core particles or core particle slurry is preferably dispersed in water (preferably deionized water). The concentration of particles in the resulting slurry may be varied, but it is preferably about 50 to 100 parts by weight of core particles per 1000 parts by weight of water. Agitation is preferably used to ensure a thorough dispersion of the core particles.
- 10 Agitation method may range from simply stirring to sonicating (i.e. using ultrasound) or combinations of methods. The slurry is also preferably heated to about 60-95°C and maintained at such elevated temperature throughout the deposition process.

- To deposit the siliceous oxide on the core particles, a siliceous oxide source (preferably sodium silicate) and an acid (preferably a mineral acid such as
- 15 hydrochloric acid or sulfuric acid) are then added to the slurry. Preferably, the slurry is maintained under agitation throughout the addition of the siliceous oxide source and the acid. The siliceous oxide source and the acid may be added in various orders of addition. Preferably, either the siliceous oxide source and acid are added simultaneously or the siliceous oxide source is added first followed by the addition of
- 20 the acid. The total proportions of siliceous oxide source and acid added to the slurry should be such that the end pH of the slurry after all additions have stopped is about 9 or less, more preferably about 6-9, most preferably about 7-8. If desired, one or more aging steps may be inserted into the process at any point during the addition. The addition of silicate and acid is preferably conducted in a manner such that
- 25 precipitation of homogeneous silica particles is substantially avoided.

- In one preferred embodiment, a quantity of aqueous sodium silicate solution (e.g., 10.7 wt.%  $\text{SiO}_2$  and 3.24 wt.%  $\text{Na}_2\text{O}$ ) containing the desired amount of silica to be deposited is added to the initial slurry of core particles slowly, e.g. over a period of at least about 45 minutes, more preferably at least about 60 minutes. An acid
- 30 solution, such as 1N HCl (or other mineral acid), is then added to the slurry until a pH of about 6-9 (preferably 7-8) is achieved. The slurry is preferably maintained under agitation at elevated temperature (preferably 60-95°C) throughout the addition of these materials. After the desired pH reduction is achieved, the acid addition is

The invention is further illustrated by the following examples, it should be understood that the invention is not limited to the specific details of the examples.

#### Example 1A

##### 5 Preparation of superparamagnetic iron oxide particles

Portions of  $\text{FeCl}_3$  and  $\text{FeCl}_2$  were dissolved in deionized water at ambient temperature to form a solution containing 3.24 wt.%  $\text{FeCl}_3$  and 1.5 wt.%  $\text{FeCl}_2$ . Under strong agitation, a 14 wt.%  $\text{NH}_4\text{OH}$  aqueous solution was quickly added to the solution until a pH of 8.5 was achieved. As a result of the ammonium hydroxide addition, a fine magnetite precipitate ( $\text{Fe}_3\text{O}_4$ ) was formed. The precipitate was allowed to settle. The most of the liquid component of the mixture was then removed by decantation. The remaining wet precipitate was then reslurried in deionized water to wash the precipitate. This wash procedure was repeated a total of three times to reduce the level of ammonium and chlorine ions. Drying of the precipitate was avoided.

#### Example 1B

##### Hydrogen peroxide treatment of superparamagnetic particles

75 g of the magnetite from Example 1A was reslurried in deionized water to achieve a 7.5 wt.% solids concentration. The resulting slurry was agitated and ultrasonically treated while 600 ml of 3 wt.%  $\text{H}_2\text{O}_2$  solution was added to the slurry. The resulting mixture was maintained in this condition for 60 minutes, and then the magnetite was separated from the solution by decantation while drying of the magnetite was avoided. The magnetite was then subjected to three washes, each using 2000 ml deionized water.

#### Example 2

75 g of wet magnetite particles prepared according to Example 1B (based on the amount of magnetite present) was then reslurried in deionized water to form a slurry having a solids content of about 7.5 wt.%. The slurry was heated to about 90°C and subjected to moderate agitation and sonication. An aqueous sodium silicate solution (10.7 wt.%  $\text{SiO}_2$  and 3.24 wt.%  $\text{Na}_2\text{O}$ ) was added to the slurry at a rate of 5 ml/min for about 43 minutes. During this addition, the sonication was

**Example 3**

75 g of wet magnetite particles prepared according to Example 1B (based on the amount of magnetite present) was then reslurried in deionized water to form a slurry (pH < 7) having a solids content of about 7.5 wt.%. While the slurry was heated to about 90°C and subjected to moderate agitation and sonication, a minor amount of an aqueous sodium silicate solution (10.7 wt.% SiO<sub>2</sub> and 3.24 wt.% Na<sub>2</sub>O) was added at a rate of 5 ml/min until a slurry pH of about 7.7 was achieved. Then, both sodium silicate solution and 1N HCl were added to the slurry simultaneously for about 43 minutes. The rate of HCl addition was adjusted to maintain the pH at 7.7 given a 5 ml/min addition rate for the sodium silicate. During this addition, the sonication was maintained for the first 15 minutes only, but the moderate agitation and 90°C temperature were maintained throughout the addition. The slurry was allowed to age for ten minutes. Then, additional amounts of the sodium silicate solution (5 ml/min) and HCl were simultaneously added for about 20 minutes while the slurry was maintained under moderate agitation at 90°C and pH=7.7. The slurry was allowed to age for about 30 minutes at 90°C under agitation, and then the particles were allowed to settle.

The particles were then subjected to the same washing and testing procedures as described in Example 2 above. The results are reported in Table 1 below. The composition of the resulting material is given in Table 3 below.

**Table 1**

Example	Surface area (m <sup>2</sup> /g)	Pore volume (<600Å diameter)	Pore volume (>600Å diameter)	Median particle size (µm)	Iron leach (ppm)
2	49	0.160	0.257	5.5	2.0
3	35	0.124	0.106	8.1	3.4

**Example 4**

125 g of wet magnetite particles prepared according to Example 1B (based on the amount of magnetite present) was then reslurried in deionized water to form a slurry having a solids content of about 7.5 wt.%. The slurry was heated to about

90°C and subjected to moderate agitation and sonication. An aqueous sodium silicate solution (10.7 wt.% SiO<sub>2</sub> and 3.24 wt.% Na<sub>2</sub>O) was added to the slurry at a rate of 8.33 ml/min for about 43 minutes. During this addition, the sonication was maintained for the first 15 minutes only, but the moderate agitation and 90°C  
5 temperature were maintained throughout the addition. The slurry was allowed to age for ten minutes. Then, additional amounts of the sodium silicate solution were added at 8.33 ml/min for about 20 minutes while the slurry was maintained under moderate agitation at 90°C. The slurry pH after the silicate addition was about 10. Under  
10 moderate agitation at 90°C, 1N HCl was then added to the slurry at a rate of about 20 ml/min until a pH of about 7.5 was achieved. The slurry was then allowed to age under agitation for about 30 minutes at 90°C, and then the particles were allowed to settle.

The particles were then subjected to the same washing and testing procedures as described in Example 2 above. The undried particles were also  
15 tested for magnetic separability using the procedure described above. These results are given in Table 2 below. The composition of the resulting material is given in Table 3 below.

#### Example 5

20 An additional sample was prepared following the same procedure as in Example 4 except that the magnetite particles were prepared according to Example 1A and that the ten minute aging step during the sodium silicate addition was eliminated.

The particles were tested as in Example 4 above. These results are given in  
25 Table 2 below. The composition of the resulting material is given in Table 3 below.

#### Example 6

An additional sample was prepared following the same procedure as in Example 2 except that the amount of magnetite particles used was 550 g. The initial  
30 slurry volume and the addition rates of sodium silicate and HCl were scaled up in proportion to the increased amount of magnetite used.

The particles were tested as in Example 4 above. These results are given in Table 2 below. The composition of the resulting material is given in Table 3 below.

**Example 7**

125 g of wet magnetite particles prepared according to Example 1A (based on the amount of magnetite present) was then reslurried in deionized water to form a slurry having a solids content of about 7.5 wt.%. The slurry was heated to about 90°C and subjected to moderate agitation and sonication. An aqueous sodium silicate solution (10.7 wt.% SiO<sub>2</sub> and 3.24 wt.% Na<sub>2</sub>O) was added to the slurry at a rate of 8.4 ml/min for about 95 minutes. During this addition, the sonication was maintained for the first 15 minutes only, but the moderate agitation and 90°C temperature were maintained throughout the addition. Under moderate agitation at 90°C, 1N HCl was then added to the slurry at a rate of about 20 ml/min until a pH of about 7.5 was achieved. The slurry was then allowed to age under agitation for about 30 minutes at 90°C, and then the particles were allowed to settle.

The particles were then subjected to the same washing and testing procedures as described in Example 4 above. These results are given in Table 2 below. The composition of the resulting material is given in Table 3 below.

**Example 8**

37.5 g of wet magnetite particles prepared according to Example 1A (based on the amount of magnetite present) was reslurried in deionized water to form a slurry having a solids content of about 7.5 wt.%. The slurry was heated to about 90°C and subjected to moderate agitation and sonication. An aqueous sodium silicate solution (10.7 wt.% SiO<sub>2</sub> and 3.24 wt.% Na<sub>2</sub>O) was added to the slurry at a rate of 5 ml/min for about 43 minutes. During this addition, the sonication, the moderate agitation and 90°C temperature were maintained. The slurry was allowed to age for ten minutes. Then, additional amounts of the sodium silicate solution were added at 5 ml/min for about 213 minutes while the slurry was maintained under moderate agitation at 90°C. Sonication was maintained for all but the last 20 minutes of the addition. Under moderate agitation at 90°C, 1N HCl was then added to the slurry at a rate of about 12 ml/min until a pH of about 7.5 was achieved. The slurry was then allowed to age under agitation for about 30 minutes at 90°C, and then the particles were allowed to settle.

## 12

maintained for the first 15 minutes only, but the moderate agitation and 90°C temperature were maintained throughout the addition. The slurry was allowed to age for ten minutes. Then, additional amounts of the sodium silicate solution were added at 5 ml/min for about 20 minutes while the slurry was maintained under moderate agitation at 90°C. The slurry pH after the silicate addition was about 10. Under moderate agitation at 90°C, 1N HCl was then added to the slurry at a rate of about 12 ml/min until a pH of about 7.5 was achieved. The slurry was allowed to age under agitation for about 30 minutes at 90°C, and then the particles were allowed to settle.

The liquid component of the settled slurry was decanted off, and the wet particles were then subjected to water washing (2000 ml deionized water) and decantation. The washing/decantation were repeated two additional times. The wet particles remaining after decantation were washed with 450 ml of 3 wt.%  $\text{NH}_4\text{Cl}$  aqueous solution for about 30 minutes followed by a further decantation. The  $\text{NH}_4\text{Cl}$  washing/decantation steps were repeated two additional times and were then followed by three additional series of deionized water washing/decantation steps.

A portion of the resulting wet product was then dried at about 110°C, and the particle size, surface area and pore volume were measured. The median particle size was determined using a Horiba light scattering device. The surface area and pore volume were determined using nitrogen BET. A portion of the dried product was tested for leachability of core metals using the test method described above. The results are given in Table 1 below. The composition of the resulting material is given in Table 3 below.

The particles were then subjected to the same washing and testing procedures as described in Example 4 above. These results are given in Table 2 below. The composition of the resulting material is given in Table 3 below.

5

Table 2

Example	Surface area (m <sup>2</sup> /g)	Pore volume (<600Å diameter)	Pore volume (>600Å diameter)	Median particle size (μm)	Iron leach (ppm)	% transmis-sion after 40 sec.	% transmis-sion after 600 sec.
4	55	0.181	0.163	5.3	2.8	87.7	98.3
5	48	0.147	0.110	5.6	3.3	96.7	100
6	56	0.188	0.142	8.0	3.7	89.9	100
7	47	0.145	0.105	6.5	2.0	99.5	100
8	6	0.020	0.017	4.7	0.7	79.2	99.4

Particle Composition

From chemical analysis of the materials resulting from the above examples 2-7, the compositions are given in Table 3 below as weight percent (dry, oxide basis).

10

Table 3

Example	SiO <sub>2</sub>	Fe <sub>2</sub> O <sub>3</sub>	Na <sub>2</sub> O
2	39.1	59.7	0.19
3	43.1	55.2	0.11
4	45.6	53.1	0.16
5	41.6	56.1	0.12
6	40.6	58.0	0.10
7	51.3	46.5	0.15
8	83.6	13.6	0.25

Claims

What is claimed is:

- 5           1. A particulate adsorbent comprising particles which consist essentially of:
- (a)   at least one core consisting essentially of a transition metal-containing component selected from the group consisting of superparamagnetic materials, low Curie Temperature materials, and mixtures thereof, and
- 10           (b)   a siliceous oxide coating on the surface of said core(s),

          wherein said coating covers the entire surface of said core(s) such that said adsorbent particles have a transition metal leach value when present as 0.33 g dried

15   adsorbent particles in 20 ml of 1N hydrochloric acid aqueous solution for 15 minutes of less than about 50 ppm metal based on the weight of said solution.

          2. The adsorbent of claim 1 wherein said transition metal is selected from Group VIII transition metals and mixtures thereof.

20           3. The adsorbent of claim 2 wherein said transition metal-containing component is selected from the group consisting of iron, iron oxides, and mixtures thereof.

25           4. The adsorbent of claim 3 wherein said transition metal-containing component consists essentially of magnetite.

          5. The adsorbent of claim 1 wherein said siliceous oxide coating contains hydroxyl groups on its outer surface.

30           6. The adsorbent of claim 5 wherein said siliceous oxide coating consists essentially of silica.

7. The adsorbent of claim 1 wherein said siliceous oxide coating contains externally accessible porosity.

5           8. The adsorbent of claim 7 wherein said particles contain at least about 0.2 ml/g pore volume measured by nitrogen BET method based on the total dry weight of said particles.

10           9. The adsorbent of claim 1 wherein said particles have a surface area of at least about 30 m<sup>2</sup>/g measured by nitrogen BET method based on the total dry weight of said particles.

15           10. The adsorbent of claim 1 wherein said core forms at least about 50 wt.% of said particles based on the combined dry basis weight of said core and said oxide coating.

20           11. The adsorbent of claim 10 wherein said core forms at least about 60 wt.% of said particles based on the combined dry basis weight of said core and said oxide coating.

          12. The adsorbent of claim 1 wherein said core comprises one or more crystals having a crystal size of about 100 nm or less.

25           13. The adsorbent of claim 12 wherein said crystal(s) have an average crystal size of about 60 nm or less.

          14. The adsorbent of claim 1 wherein said adsorbent particles have an average particle size of about 1-15  $\mu$ m.

30           15. The adsorbent of claim 14 wherein said adsorbent particles have an average particle size of about 3-10  $\mu$ m.

16. The adsorbent of claim 1 wherein said core is a superparamagnetic material at room temperature.

17. The adsorbent of claim 8 wherein said particles contain at least about 0.2 ml/g porosity in pores having a diameter of 60 nm or greater as measured by nitrogen BET method.

18. The adsorbent of claim 16 wherein said superparamagnetic material has a remanent magnetism level of about 10 emu/g or less.

10

19. The adsorbent of claim 1 wherein said core(s) consists essentially of a material having a Curie Temperature of about -50°C. to 100°C.

20. A method of making adsorbent particles wherein said particles comprise (a) one or more cores consisting essentially of a transition metal-containing component selected from the group consisting of superparamagnetic materials, low Curie Temperature materials and mixtures thereof, and (b) a siliceous oxide coating on the surface of said core(s), said method comprising:

- (i) forming an aqueous slurry of said cores,
- (ii) adding a siliceous oxide precursor to said slurry,
- (iii) adding an aqueous acid solution to said slurry whereby the pH of said slurry is reduced to about 6-9, and
- (iv) recovering said adsorbent particles from said slurry resulting from step (iii).

25

21. The method of claim 20 wherein said slurry is aged for at least about 15 minutes between steps (iii) and (iv).

22. The method of claim 20 wherein step (i) comprises precipitating said cores in an aqueous medium.

30

23. The method of claim 22 wherein steps (ii) and (iii) are conducted at a temperature of about 60 - 95°C.

24. The method of claim 20 further comprising:

- (v) washing said recovered particles with water, and
- (vi) further washing said recovered particles with a dilute ammonium chloride solution.

5

25. The method of claim 20 wherein said siliceous oxide source comprises an aqueous solution of alkali metal silicate.

26. The method of claim 20 wherein said acid comprises a mineral acid.

10

27. The method of claim 20 wherein said cores consist essentially of magnetite.

28. The method of claim 20 wherein steps (ii) and (iii) are conducted simultaneously.

15

# INTERNATIONAL SEARCH REPORT

Int. Appl. No.  
PCT/US 98/00533

**A. CLASSIFICATION OF SUBJECT MATTER**  
IPC 6 801J20/28 801J20/10

According to International Patent Classification (IPC) or to both national classification and IPC

**B. FIELDS SEARCHED**

Minimum documentation searched (classification system followed by classification symbols)

IPC 6 801J

Documentation searched other than minimum documentation to the extent that such documents are included in the fields searched

Electronic data base consulted during the international search (name of data base and, where practical, search terms used)

**C. DOCUMENTS CONSIDERED TO BE RELEVANT**

Category	Citation of document, with indication, where appropriate, of the relevant passages	Relevant to claim No.
A	US 5 039 559 A (SANG) 13 August 1991 see column 2, line 45 - column 4, line 24	1-4, 10-16,19
P,A	EP 0 757 106 A (TOYO BOSEKI) 5 February 1997 see page 4, line 10 - page 5, line 26	1-4, 6-11, 14-16, 20,25,27
A	DE 43 07 262 A (BERGEMANN) 8 September 1994 see column 1, line 21 - column 2, line 2	1,5,20, 22,25-27
A	WO 96 03653 A (SILICA GEL) 8 February 1996 see page 15-21; claims 1,6,12	1



Further documents are listed in the continuation of box C.



Patent family members are listed in annex.

**\* Special categories of cited documents:**

- "A" document defining the general state of the art which is not considered to be of particular relevance
- "E" earlier document but published on or after the international filing date
- "L" document which may throw doubts on priority claims or which is cited to establish the publication date of another citation or other special reason (see specification)
- "O" document referring to an oral disclosure, use, exhibition or other means
- "P" document published prior to the international filing date but later than the priority date claimed

- "T" later document published after the international filing date or priority date and not in conflict with the application but cited to underscore the principle or theory underlying the invention
- "X" document of particular relevance; the claimed invention cannot be considered novel or cannot be considered to involve an inventive step when the document is taken alone
- "Y" document of particular relevance; the claimed invention cannot be considered to involve an inventive step when the document is combined with one or more other such documents, such combination being obvious to a person skilled in the art.
- "Z" document member of the same patent family

Date of the actual completion of the international search

8 May 1998

Date of mailing of the international search report

18/05/1998

Name and mailing address of the ISA  
European Patent Office, P.O. Box 2916 Paternoster 2  
NL - 2280 HV Rijswijk  
Tel. (+31-70) 340-6040, Tx. 31 651 apo nl,  
Fax: (+31-70) 340-3916

Authorized officer

Wendling, J-P

# INTERN. 'ONAL SEARCH REPORT

Information on patent family members

Info. Int. Application No

PCT/US 98/00533

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
US 5039559 A	13-08-1991	AT 117829 T	15-02-1995
		DE 68920778 D	09-03-1995
		DE 68920778 T	18-05-1995
		EP 0343934 A	29-11-1989
		ES 2066851 T	16-03-1995
		JP 2038318 A	07-02-1990
		US 5662824 A	02-09-1997
EP 757106 A	05-02-1997	JP 9019292 A	21-01-1997
DE 4307262 A	08-09-1994	NONE	
WO 9603653 A	08-02-1996	DE 4427821 A	01-02-1996
		EP 0772776 A	14-05-1997
		JP 10503281 T	24-03-1998